# The ELi<sup>™</sup> Process – Greener Lithium for Battery Materials

By

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## Lithium Hydroxide for the EV Era

### The old days – A Niche Material:

- · Small volume production, limited installed capacity
- · Limited quantities of product suitable for battery cathode
- Use of bulk industrial chemical reagents for two step chemical conversion via lithium carbonate
- Long distance haulage of bulk materials was relatively insignificant
- The relatively high carbon footprint was not a factor
- Relatively high cost through "double conversion"

### The EV Era Needs:

- Large volume production
- Consistent and high purity
- Small carbon footprint
- Lower cost production
  - Protection for producers during lower price phases
  - Lower cost for cathode and battery producers under pressure to reduce battery cost
- · Production from both brines and hard rock sources to meet demand growth

## **Lithium Hydroxide Production**

### The old days – Chemical Conversion:

### Hard Rock

•  $\text{Li}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 = \text{Li}_2\text{CO}_3 + \text{Na}_2\text{SO}_4$ 

energy-intensive production

• Li<sub>2</sub>CO<sub>3</sub> + Ca(OH)<sub>2</sub> = 2 LiOH + CaCO<sub>3</sub>

### Brine

- $2\text{LiCl} + \text{Na}_2\text{CO}_3 = \text{Li}_2\text{CO}_3 + 2\text{NaCl}$
- Li<sub>2</sub>CO<sub>3</sub> + Ca(OH)<sub>2</sub> = 2 LiOH + CaCO<sub>3</sub>

The mode	rnized o	ld days:			
Hard rock					

•  $Li_2SO_4 + 2NaOH = 2 LiOH + Na_2SO_4$ 

Low market value by-product with energy-intensive production

Low market value by-product

Bulk residue, disposal cost

Bulk residue, disposal cost

with

### The Modern ELi™ Process:

 $2\text{LiCl} + 2\text{H}_2\text{O} + \text{e} = 2\text{LiOH} + \text{Cl}_2 + \text{H}_2$  (synthesised into 32% HCl)

- Suitable to convert spodumene concentrates and sub-surface lithium brines
- Low carbon electricity
  - Has increasing availability
  - Further reduction of carbon footprint
- No bulk industrial chemical reagents
  - Minimises heavy transport costs
  - Reduces carbon footprint associated with production of reagents and transport of reagents
  - Reduced pressure on the reagent supply chains
- Minimised waste materials and disposal

## ELi<sup>™</sup> – A New Technology Solution to an Old Problem

Flowsheets for processing spodumene and for processing sub-surface brines



### Where ELi<sup>™</sup> fits into the Lithium Brine Production Chain

Combined DLE and ELi<sup>™</sup> Process for Lithium Production



## **The Starting Point**

- Brine compositions are highly variable
  - Typical impurity concentrations and ratios vary salar to salar
  - Vary between brine and spodumene sources
- Purify to a higher degree than conventional lithium processes
- Minimise lithium losses associated with impurity removal

#### Typical Sub-surface Brine after pre-treatments

(mg/L)	Li	Са	Mg	В	Sr	Si
Before	33,000-49,000	3,400-6,500	8,600-13,000	3,500-5,400	30-40	20-40

### The "Aha" Moment

Membrane electrolysis directly converts lithium chloride to hydroxide and eliminates transport & use of bulk reagents

Process	ELi™	Conventional Processes		
Bulk Lime	No	Yes		
Bulk Soda Ash	Minimal	Yes		
Power	Yes (renewable)	Yes		
Gas/oil	Yes	Yes		

## The Enabler – Ultrapure LiCl Feed to Electrolysis

- Successful brine purification pilot enables efficient electrolysis
  - The higher purity the solution, the higher the product purity and the lower the power cost
  - Flowsheet successfully piloted
  - Previous electrolysis tests were successful on less pure feed
- Next piloting phases are long duration electrolysis tests and LHM crystallisation
  - Electrolysis tests are evaluating membrane durability and cell hydrodynamics

(mg/L)	Li	Са	Mg	В	Sr	Si
Before	62,654	6,260	15,792	7,141	71	52
After (ave)	45,100	< 0.9	< 0.09	< 0.4	< 0.002	1
% removed*	n/a	> 99.98	> 99.99	> 99.99	> 99.99	> 97

## **ELi™ Membrane Electrolysis**



# **ELi™ Membrane Electrolysis Equipment**



Source: Denora

### Low Opex is Compatible with Periods of Low Market Prices





#### Estimated Opex Comparison (Conversion to LHM)

Source: ALB, E3 Li, PLL, Livent, Management estimated, Class 3 ECS, Benchmark Mineral Intelligence.

Source: Benchmark

## **Commercialisation – Further Piloting and Scale-Up**

#### Disciplined, gated approach to reaching commercial operations

#### Process batch pilot evaluation:

- Performed on a Sth American brine eluate
- Pre-concentration completed and characterized
- Purification completed
- Long-duration electrolysis commencing
- · Product crystallisation from catholyte to follow
- Next step is sub-commercial scale continuous pilot

#### Plant design and estimation:

- Vendor packages were integrated into Engineering Cost Study to AACE Class 3 (+/- 15% accuracy)
- Plant layouts and PFDs completed, brownfields site
- Partial completion of P+IDs
- Capex and Opex estimates completed
- Process LCA on gate-to-gate basis completed, in peer review



### Result – Lowest Opex Conversion & Potentially Lowest CO<sub>2</sub>



Estimated Opex Comparison (Conversion to LHM)

#### Estimated Capital Intensity (Conversion to LHM)

US\$ Capex/t Annual LHM Capacity



Source: ALB, E3 Li, PLL, Livent, Management estimated, Class 3 ECS, Benchmark Mineral Intelligence.



Source: SQM 2024

## **Sustainable Competitive Advantages for Plant Operators**

Commercial operating licence for highly efficient, lithium hydroxide plant utilising our patented processing flowsheet



No bulk reagents ensures essential lithium produced at substantial discount to conventional process



Replacement of carbonheavy traditional bulk reagents



Reduces the lithium carbon footprint for EV batteries

- Commercial deployment of ELi<sup>™</sup>
  - · Licensing business model, to producers, for royalty payments
  - Partnership model with producers
- RAM holds 17 granted patents and 14 pending patents worldwide
  - 4 families of patents covering brine and hard rock processing
- Major lithium resource production regions
  - Incl USA, Argentina, Chile, Australia